

CHAPTER 1: PARTICLES OF MATTER

Classifications of Matter

- Matter = mass + volume
- Compositions of matter- based on the kinds and amounts of substances that compose it
 - o Pure substances – made of only one kind of atom or molecule
 - Elements ex Na
 - Compounds ex H₂O
 - o Mixtures – made of 2+ kinds of atoms or molecules
 - Homogeneous – every sample identical in composition
 - Heterogeneous - every sample is different
- Physical properties – can be noted without changing its composition ex density
 - o Physical change = changes of state – you aren't making anything new
 - Ex. melting ice
- Chemical properties – can only be observed by changing its composition ex flammability
 - o Chemical change = formation of a new substance
- Atom – The smallest particle of an element that retains the chemical characteristics of that element
- Molecule – of two or more atoms that are joined together by chemical bonds
- Chemical formula – Notation used to represent elements and compounds ex) Water H₂O
- Chemical equation - Use chemical formulas to show substances involved in a reaction and what products they make

Units of Measurement

- Make sure you can interconvert between different units using the metric system
- Overall unit conversion “formula” – very useful in stoichiometry
 - o $given\ units \times \frac{desired\ units}{given\ units} = desired\ units$
 - If the question has multiple unit conversions, make a plan
 - Ex) grams of x → moles of x → moles of Y → grams of Y
- $1ml = 1cm^3$
- Always check units!

SIG FIGS

- 1) All non-zero digits are significant ex 52 = 2 sigfigs
- 2) zeros in the middle are significant ex 502 = 3 sigfigs
- 3) zeros added to the front of the number are NOT significant ex 0000000052 = 2 sigfigs
- 4) zeros at the end may or may not be significant
 - after a decimal point, they are significant ex 52.0 = 3 sigfigs
 - before a written decimal point are ambiguous and should be avoided by using scientific notation ex. 520 = ambiguous

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