Activity 2

Directions. Classify the following unbalanced chemical equations according to the six types of chemical reactions. Write your answers in your notebook/ on a separate sheet of paper.

- 1. $AgNO_3 + Cu \rightarrow Cu(NO_3)_2 + Ag$
- P₄ + O₂ → P₂O₅
- HBr + Al(OH)₃ → AlBr₃ + H₂O
- 4. $O_3 \rightarrow O' + O_2$
- 5. $C_3H_6O + O_2 \rightarrow CO_2 + H_2O$

- HBr + KOH → KBr + H₂O
- H₂ + N₂ → NH₃
- 8. $Hg_2O \rightarrow Hg + O_2$
- Cl₂ + NaBr → NaCl + Br
- 10. $CH_4 + O_2 \rightarrow CO_2 + H_2O$
- 11. $SO_2 + O_2 \rightarrow SO_3$
- 12. HgO → Hg + O₂
- 13. $P + O_2 \rightarrow P_4O_{10}$
- 14. KClO₄ → KCl+ O₂
- 15. $Fe_2O_3 + C \rightarrow Fe + CO_2$
- 16. $H_2 + AgNO_3 \rightarrow Ag + HNO_3$
- CaCO₃ + HCl → CaCl₂ + H₂CO₃
- 18. $Cu + AgNO_3 \rightarrow Ag + Cu(NO_3)_2$
- 19. $Mg + O_2 \rightarrow MgO + energy$
- HCl + NaOH → NaCl + H₂O

Classifying Chemical Reactions Study Guide

Leo J. Malone

Classifying Chemical Reactions Study Guide:

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chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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