

Gases | Study Guide

Test Format

Multiple Choice
Gas Law Problems
Gas Stoich Problems
Short Answer

Equations & Conversions

$$P = \frac{F}{A}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$PV = nRT$$

$$R = 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

$$1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$$

$$\frac{\text{Rate of Effusion of L}}{\text{Rate of Effusion of H}} = \sqrt{\frac{\text{Molar Mass of H}}{\text{Molar Mass of L}}}$$

$$K = ^\circ\text{C} + 273 \quad 0.001 \text{ kL} = 1 \text{ L} = 1,000 \text{ mL}$$

$$\text{At STP, 1 mole} = 22.4 \text{ L} \quad \text{KHDSdcm}$$

1. A box weighs 180 lbs and has a base that is 6 inches by 9 inches. What is the pressure this box exerts on the surface it is sitting on?

$$P = \frac{F}{A} \quad \frac{180 \text{ lbs}}{54 \text{ in}^2} = \boxed{3.3 \text{ psi}}$$

$$54 \text{ in}^2 = \text{area}$$

2. What is atmospheric pressure?

The pressure exerted by the collisions of gases in the air surrounding Earth (N_2 , O_2 , etc).

3. Convert 1.088 atm to:

a. mm Hg $\frac{1.088 \text{ atm}}{1} \cdot \frac{760 \text{ mmHg}}{1 \text{ atm}} = \boxed{826.9 \text{ mmHg}}$

b. kPa $\frac{1.088 \text{ atm}}{1} \cdot \frac{101.3 \text{ kPa}}{1 \text{ atm}} = \boxed{110.2 \text{ kPa}}$

4. What is a barometer and who invented it?

A tool that measures room/air pressure. Invented by Torricelli.



5. How are temperature and kinetic energy related?

Direct. As temp $\uparrow$, KE $\uparrow$, molecules move faster.
As temp $\downarrow$, KE $\downarrow$, molecules move slower.

6. Briefly describe the five assumptions of the kinetic-molecular theory.

Gases have a lot of tiny particles that are always in motion that is random, fast, and continuous. Particles are far apart without any forces of attraction, so their collisions are elastic $\rightarrow$ bouncing off of each other/walls. The speed of the molecules change with temperature.

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