

Lab - Measuring Mass: A Means of Counting

Adapted from Prentice Hall Chemistry, Small Scale Chemistry Laboratory Manual

OBJECTIVES

- Measure masses of common compounds, objects, and minerals.
- Calculate moles and atoms from experimental masses.

Many everyday items are counted in groups in order to make them more manageable. Eggs are counted in dozens; sheets of paper are counted in reams. Since atoms and molecules are so minuscule, it is necessary to count them in groups much larger than 12 or 500. Instead, they are counted in groups called moles. The mole is a group of 6.02×10^{23} particles, or Avogadro's number of particles.

Counting the individual items in a group as large as the mole is impossible. As a result, a different method is used to count a mole of molecules - counting by weighing. The same process is used to count aluminum cans when you take a bag to the recycling center. A conversion factor (weight per can) is used to convert the weight of the bag into the number of cans in the bag. In the same way, the molar mass of an element or compound is used to convert the mass of a sample into the number of atoms or molecules in that sample.

In this lab, you will perform molar conversions for small samples of everyday substances.

PRELAB QUESTIONS

1. How is a mole like a dozen?
2. How many pieces or particles are in 1 mole?
3. How can we count how many particles a substance contains?
4. What is the conversion factor to convert the mass of a sample to the number of moles?

PURPOSE

In this experiment, you will measure the masses of samples of various common compounds such as water, salt, and sugar. You will use your results as a means of counting the atoms, ions, and molecules in your samples. You will extend your technique to common objects that you can consider to be pure substances, such as glass marbles, pieces of chalk, and polystyrene peanuts. Finally, you will measure the masses of various mineral samples and use your results to find the number of atoms in each.

SAFETY

- Behave in a way that is consistent with a safe laboratory.
- Don't chew gum, drink, or eat in the laboratory. Never taste a chemical in the laboratory.

MATERIALS

sodium chloride (NaCl)

sucrose ($C_{12}H_{22}O_{11}$)

chalk ($CaCO_3$)

other common minerals specified by your teacher

plastic spoons

balance

water (H_2O)

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