

Collision Theory - answers

Fill in the blanks using each of the words once.

energy	collide	collisions	concentration	collision theory
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Particles can only react if they **collide** with enough **energy** for the reaction to take place. This is called the **collision theory**. There are four factors that can change the rate of a reaction; temperature, **concentration**, surface area and the use of a suitable **catalyst**.

Fill in the blanks in the text and complete the diagrams below using the words below.

moderate	fast	faster	energy	surface area
faster	faster	slow	low concentration	fast
faster	more often	collisions	successful	slow
particles	faster	successful potential	high concentration	large surface area

Temperature

Increasing the temperature will cause the particles to move **fast** with more energy. They will therefore collide **more often** and with greater **energy**. These two things mean there are more successful collisions per second and therefore a **faster** rate of reaction.

Concentration

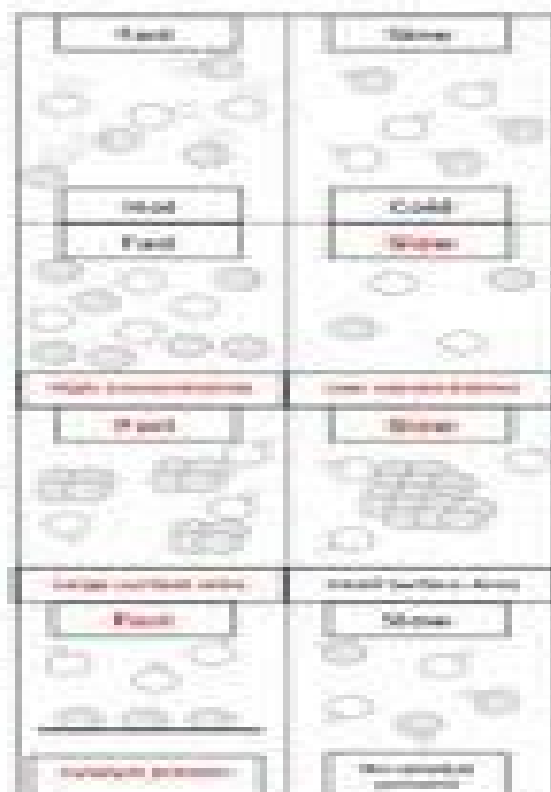
Increasing the concentration of a reactant simply means there are more **particles** which may collide and so react. More collisions means a **faster** reaction.

Surface Area

Using a powder instead of a lump means **surface area** is greater, which means a greater area of reactant is exposed and so available for a collision. More collisions means a **faster** reaction.

Catalysts

Use of a suitable catalyst means that the particles may react even if they collide with only **moderate** energy. This means more **successful** collisions are likely. Some catalysts work because one of the particles is fixed to a surface. This makes the chance of a **collision** more likely. More collisions means a **faster** reaction.



Collision Theory Worksheet 161 Answer Key

Shasha Hu

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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