

PROBLEMS WITH CALCULATIONS

1. Nitrogen gas and hydrogen gas combine according to the balanced equation below:



The diagram below shows an initial mixture of nitrogen gas (blue molecules) and hydrogen gas (red spheres) which react to form ammonia (blue and red spheres). Before the reaction, the total number of molecules is 10. After the reaction, the total number of molecules is 6.



Questions:

1. How many molecules of nitrogen gas and hydrogen gas are present in the initial mixture? How many molecules of ammonia gas are present in the final mixture? How many molecules of hydrogen gas are present in the final mixture?

2. How many grams of nitrogen gas and hydrogen gas are present in the initial mixture?

- a. How many grams of nitrogen gas and hydrogen gas are present in the initial mixture?

$$\begin{aligned} \text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) &\longrightarrow 2\text{NH}_3(\text{g}) \\ 28.0 \text{ g N}_2 &+ 6.0 \text{ g H}_2 \longrightarrow 34.0 \text{ g NH}_3 \end{aligned}$$

- b. What is the limiting reagent? (showing work)

$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$$

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$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$$

- c. What is the percent yield of the reaction? (showing work)

$$\frac{34.0 \text{ g}}{34.0 \text{ g}} \times 100\%$$

$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$$

Ccsf Chem 40 Lab Manual

Jin-Ying Zhang



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