

Name _____

Section _____

Date _____

Post-Laboratory Questions

(Use the spaces provided for the answers and additional paper if necessary.)

1. A student doing the experiment in this module extended the study of the $[\text{CoCl}_4]^{2-}$ ion/ $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ion equilibrium. In one test tube, the student added silver nitrate (AgNO_3) to a blue equilibrium mixture prepared from CoCl_2 . The reaction mixture became pink and cloudy, and a white precipitate settled out leaving a clear, pink solution. The student identified the precipitate as silver chloride, AgCl .

(1) Briefly explain how these observations are consistent with Le Châtelier's principle.

In a second test, the student placed a test tube containing a pink equilibrium mixture in a hot-water bath. The solution turned blue. When the student removed the test tube from the hot-water bath and placed it in an ice-water bath, the solution turned pink.

(2) Is the forward reaction in the $[\text{CoCl}_4]^{2-}$ ion/ $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ion equilibrium exothermic or endothermic?

(3) Briefly explain how the student's observations support your answer to (2).

(4) Write the net ionic equation for this equilibrium, including heat.

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L Towne



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Metals Abstracts ,1993

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