CHAPTER 7 REVIEW

Chemical Formulas and Chemical Compounds

SECTION 7-1

SHORT ANSWER Answer the following questions in the space provided.

- In a Stock name such as iron(III) sulfate, the roman numeral tells us ______
 - (a) how many atoms of Fe are in one formula unit
 - (b) how many sulfate ions can be attached to the iron atom.
 - (c) the charge on each Fe ion
 - (d) the total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to ______
 - (a) change the number of moles represented by the formula
 - (b) change the charges on the other ions in the compound
 - (c) change the formula so that it no longer represents the compound it previously represented
 - (d) have no effect on the formula
- The explosive TNT has the molecular formula C₂H₃(NO₂)₃.
 - 4 elements a, How many elements make up this compound?
 - 6 oxygen atoms b. How many oxygen atoms are present in one molecule of $C_2H_3(NO_2)_3$?
 - 21 atoms c. How many atoms in total are present in one molecule of C₇H₅(NO₂)₃?
 - 4.2 × 10²⁴ atoms d. How many atoms are present in a sample of 2 × 10²³ molecules of C₇H₂(NO₂)₃?
- 4. How many atoms are present in each of these formula units?
 - 11 atoms a, Ca(HCO₃)₂
 - 45 atoms b. C₁₂H₂₂O₁₁
 - 10 atoms c. Fe(ClO₂)₃
 - 9 atoms d. Fe(ClO₃)₂
- 5. N₂Q₅ a. What is the formula for the compound dinitrogen pentoxide?
 - carbon(IV) sulfide b. What is the Stock name for the covalent compound CS₂?
 - H₂SO₃ c. What is the formula for sulfurous acid?
 - phosphoric acid d. What is the name for the acid H₃PO₄?

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Jill Meryl Levy

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 - Chapter 2: Essential Elements of Chemical Names And Formulas Study Guide Answers
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 - Chapter 4: Chemical Names And Formulas Study Guide Answers in Specific Contexts
 - ∘ Chapter 5: Conclusion
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